Name: Date:
Period: Grahams Law Worksheet

1. If equal amounts of helium and argon are placed in a porous container and allowed to escape, which gas will escape faster and how much faster?

2. What is the molecular weight of a gas which diffuses 1/50 as fast as hydrogen?

3. Two porous containers are filled with hydrogen and neon respectively. Under identical conditions, 2/3 of the hydrogen escapes in 6 hours. How long will it take for half the neon to escape?

4. If the density of hydrogen is 0.090 g/L and its rate of diffusion is 6 times that of chlorine, what is the density of chlorine?

5. How much faster does hydrogen escape through a porous container than sulfur dioxide?